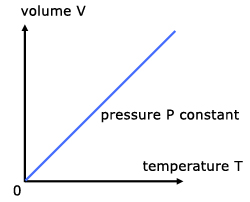
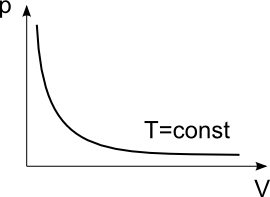
1. **State** the law that corresponds to the following graphs:

1. **Order** the following elements by increasing electronegativity: O, Li, Fe, F.
2. **State** the type of bonding between the atoms of a substance that has the following properties: low melting point, non-soluble in water but soluble in organic solvents, non-conductor.
3. Formulation

a) Barium oxide: f) CH4:

b) Phosphoric acid: g) HNO2:

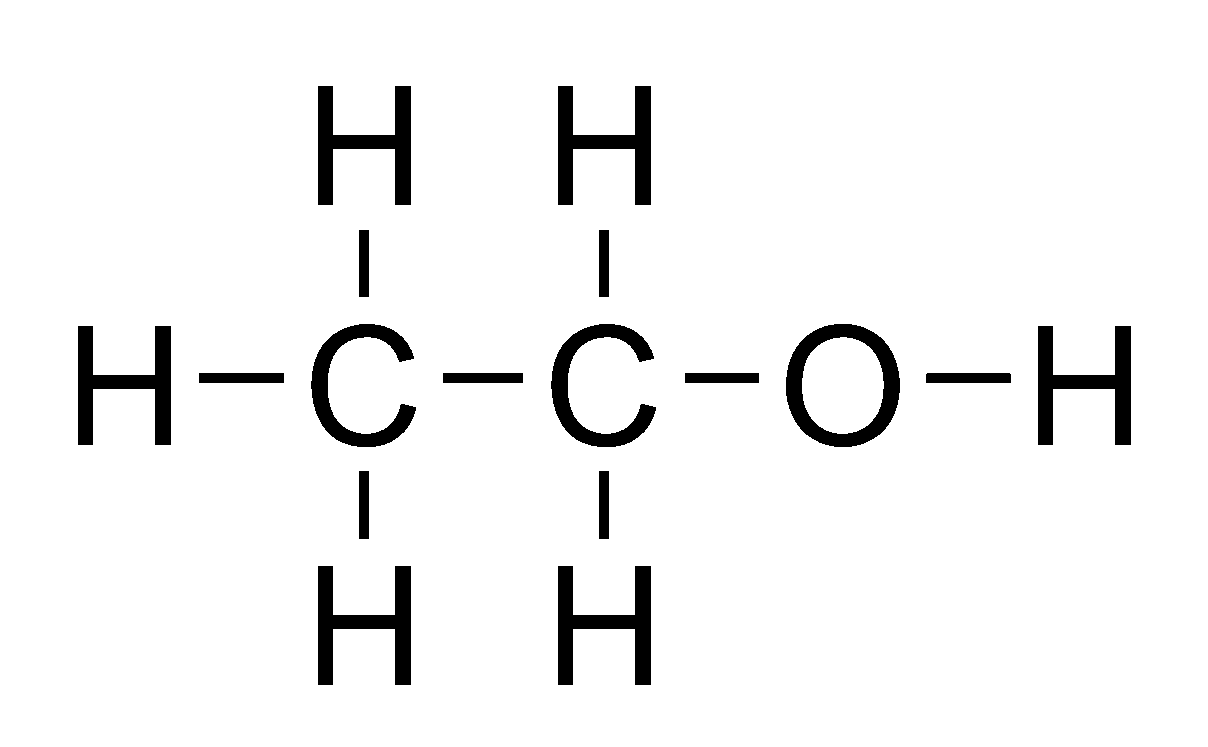
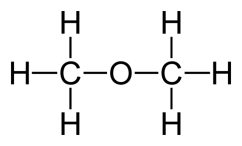
c) Ammonium chloride: h) MgCrO4:

d) Tin(IV) hydroxide: i) MnO2:

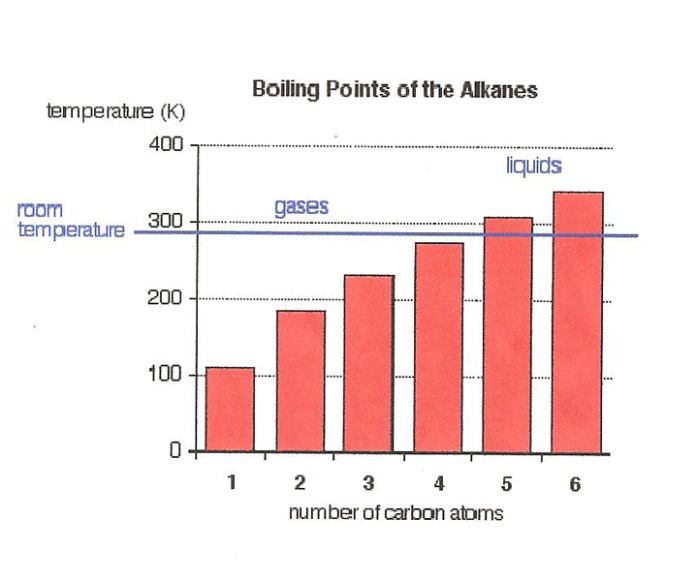
e) Cobalt(II) nitrate: j) H2S:

1. **Explain** why Bohr’s model can explain the atomic spectra, but previous models cannot.
2. Using the following table, **indicate, giving the reasons,** two elements that would bond using ionic bonding, two elements that would use covalent bonding, and two elements that would use metallic bonding.

|  |  |  |  |  |
| --- | --- | --- | --- | --- |
| Element | O | S | Zn | Zr |
| Electronegativity | 3.5 | 2.6 | 1.6 | 1.4 |

1. **Explain** the changes of state on the basis of the kinetic theory.
2. We have 200 cm3 of a gas, initially at 27 ºC. **Calculate** the final volume if we increase the temperature at 87 ºC while maintaining constant pressure.
3. These are the structures of ethanol and dimethyl ether, two organic compounds with exactly the same molecular formula, and so, the same molecular mass. However, ethanol´s boiling point is 78 ºC, while dimethyl ether´s is -24 ºC. **Explain** this fact.

Ethanol Diethyl ether

1. **Explain** metallic bonding and **state** the properties of metallic substances.
2. **Determine** the molar mass of an unknown gas that has a volume of 4.1 L at a temperature of 27 °C, a pressure of 1.5 atm, and a mass of 11 g. Data: R=0.082 atm·L/K·mol
3.  Alkanes are compounds that contain only carbon and hydrogen atoms.

The graph to the right shows the boiling points of alkanes with different numbers of carbon atoms.

**State** and **explain** this trend.